

KSCE TOP PREDICTION MASTER CYCLE 3

233/3 CHEMISTRY PRACTICAL – END OF TERM 1

CONFIDENTIAL INSTRUCTIONS

In addition to the laboratory fittings each candidate requires the following:

- About 60cm³ of solution P
- About 80cm³ of solution Q
- Accurately weighed 1.5g of solid R in a stoppered container.
- Solid H- about 0.5g in a stoppered container
- Solid G – about 0.5g in a stoppered container
- A burette
- A pipette and pipette filler
- Two conical flasks
- A 100ml plastic beaker
- A 250ml beaker
- Thermometer
- One stop watch
- 5 test-tubes
- One boiling tube
- 100ml measuring cylinder
- A small piece of aluminium foil
- A spatula
- Red and blue litmus paper
- Universal indicator paper and PH chart
- Test tube holder
- Filter paper
- Filter funnel

Access to

- 2M sodium hydroxide solution
- 2M ammonia solution
- Distilled water
- Phenolphthalein indicator
- Source of heat
- Acidified potassium dichromate (vi)
- Acidified potassium manganate (vii)
- Lead (ii) nitrate solution
- Potassium iodide solution

Preparations

- Solution P is 2M dilute hydrochloric acid
- Solution Q is 0.6M sodium hydroxide solution
- Solid R is accurately weighed 1.5g of sodium hydrogen carbonate solid
- Solid H is lead (ii) nitrate solid
- Solid G benzoic acid