

NAME.....CLASS.....ADM.....SIGN...

**CHEMISTRY**

**FORM FOUR**

**(THEORY) PAPER ONE**

**TIME: 2 HOURS**

# **KCSE TOP PREDICTION MASTER CYCLE 1**

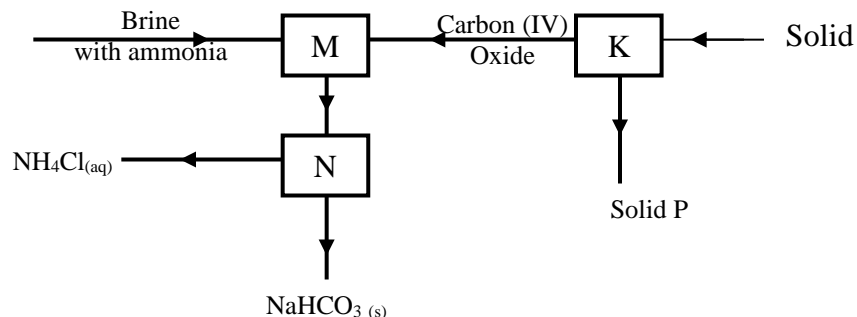
## **INSTRUCTIONS TO STUDENTS:**

- Write your **name** and **adm** in the spaces provided **above**.
- **Sign** and write the **date** of examination in the spaces provided **above**.
- Answer **ALL** the questions in the spaces provided.
- All working **must be** clearly shown where necessary
- Candidates should check the question paper to ascertain that all the pages are printed as indicated and that no questions are missing
- Candidates should answer the questions in English.

### **For Examiner's Use Only**

<b>Questions</b>	<b>Maximum Score</b>	<b>Candidate's Score</b>
<b>1 – 29</b>	<b>80</b>	

1. The diagram below shows part of Solvay process.



(a) Name solid P ( 1 Mark)

.....

(b) State the process taking place in chamber N.

( 1mark)

.....

(c) State two uses of calcium chloride which is a by-product in this process.( 1 mark)

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2.  $100\text{cm}^3$  of methane gas diffused through a porous partition in 40 seconds. How long would it take  $90\text{cm}^3$  of ozone gas to diffuse through the same partition? C = 12, H = 1, O = 16 (3marks)

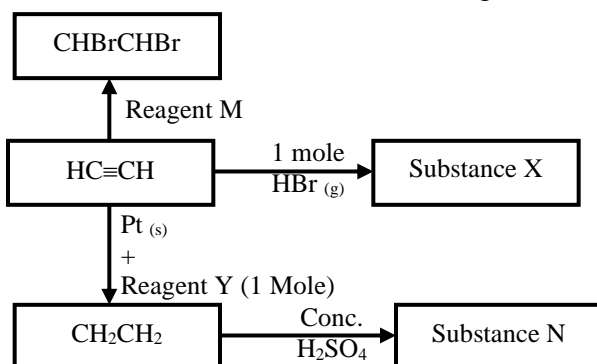
3. Ammonia is produced in large scale by Haber process.

(i) Write an equation for the formation of ammonia gas.

(1 mark)

(ii) State **two** optimum conditions for obtaining a high yield of ammonia in the process.  
(2 marks)

4. The scheme below shows some reactions starting with ethyne. Study it and answer the questions that follow.



(a) Name substance X and N (1mark)

(b) Name reagent M

(1 Mark)

(c) Ethene undergoes polymerization to form a polymer. Give an equation for the reaction and name the product.

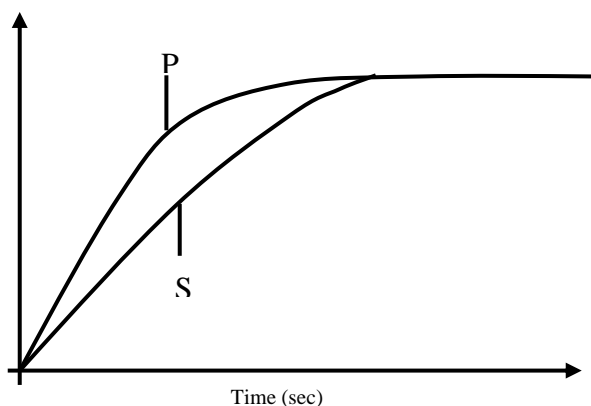
(i) Equation;

(1 mark)

(ii) Name:

(1mark)

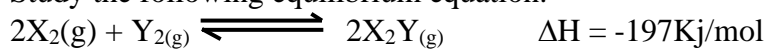
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5. The curves below represent the volume of carbon (IV) oxide gas evolved once 2M(concentrated) hydrochloric acid was reacted with 100g of powdered calcium carbonate and also when 1M concentrated hydrochloric acid was reacted with the same quantity of carbonate.



- (i) Which of the two curves represents the reaction of 2M concentrated HCl with powdered calcium carbonate. Give a reason. (2 marks)

- .....
- .....
- .....
- (ii) Why do the two curves flatten at the same level of production of CO<sub>2</sub> (1 mark)

6. Study the following equilibrium equation.



Suggest two ways of increasing the yield of X<sub>2</sub>Y.

(1 mark)

7. The table below gives some elements in the periodic table. Use it to answer the questions that follow. The letters do not represent the actual symbols of the elements.

Element	A	B	C	D	E
Atomic number	12	13	14	15	16

Which of the above letters represent:

(a) A metallic element which forms ions with the smallest ionic radius? Explain(2 marks)

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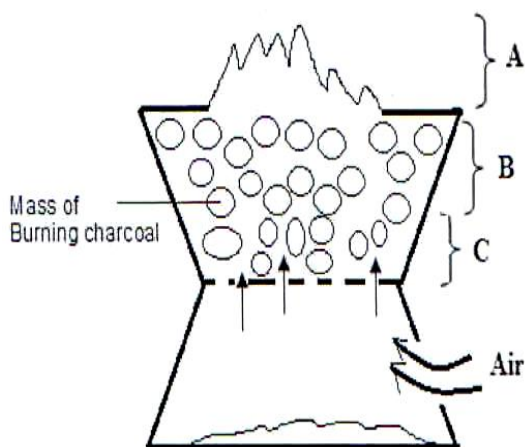
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(b) A non metallic element with the largest atomic size? Explain. (1 mark)

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8. The diagram below shows a burning jiko. Study it and answer the questions that follow.



(a) Write the equation for the reaction taking place in region A. (1 Mark)

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(b) Name the gas produced at region B. (1 Mark)

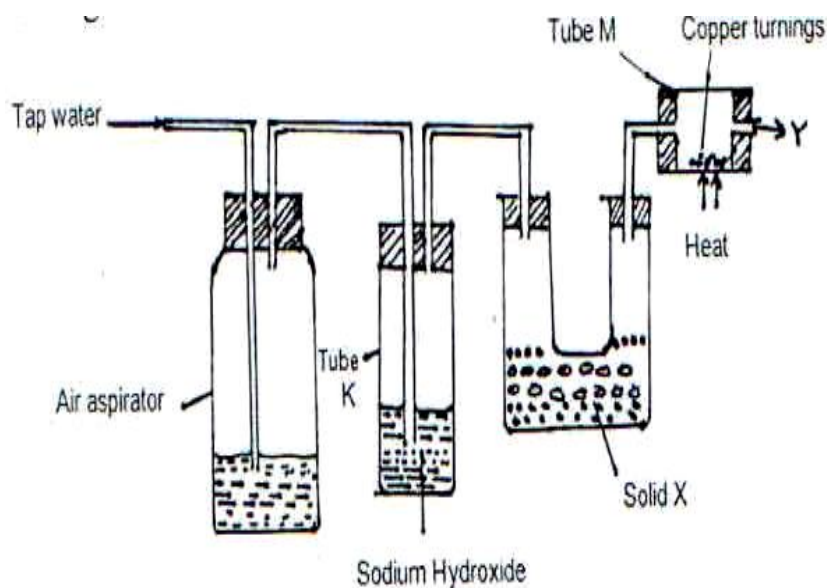
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(c) State ONE use of the gas named in (b) above.

(1 Mark)

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9. Study the diagram below and answer the questions that follow.



(i) What is the purpose of passing tap water through the air aspirator? (1 Mark)

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(ii) State and explain the observation that would be made in tube M after sometime. (1 Mark)

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10. 15g of sodium chloride was dissolved in 120cm<sup>3</sup> of distilled water. Calculate the concentration of the resulting solution in moles per litre. (Na = 23, Cl = 35.5) **(3Marks)**

11. (a) State Boyle's Law.

**(1 Mark)**

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.....

(c) The volume of a gas at 30<sup>0</sup>C and 780mmHg is 400cm<sup>3</sup>. What will be its volume at 50<sup>0</sup>C at 600 mmHg. **(3marks)**

12. Sulphur exhibits allotropy.

(a) What is allotropy?

**(1 Mark)**

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.....

(b) Name the **two** allotropes of sulphur.

**(2 Marks)**

(c) Sulphur powder was placed in a deflagrating spoon and heated on a Bunsen Burner.

(i) State the observation made. *(1 Mark)*

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(ii) The product obtained was dissolved in water. Comment on the PH of the solution formed.*(1 Mark)*

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13. 0.318g of an oxide of metal M was completely reduced by hydrogen gas to 0.254g of metal. Calculate empirical formula of the metal oxide. (M = 63.5, O = 16) *(3 Marks)*

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14. Given the following reagents: Solid sodium Carbonate, water, solid Lead (II) nitrate. Describe how a sample of Lead (II) Carbonate can be prepared in the laboratory.*(3 Marks)*

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15. Volume of liquids can be measured using a pipette; measuring cylinder or burette. Explain which one would be best for measuring 29.1cm<sup>3</sup> of liquid. *(1 Mark)*

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16. Study the information in the table and answer the questions below.

Substance	Solubility g/100g water
V	126
W	2

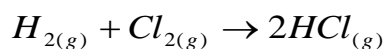
Describe how a solid sample of substance **V** could be obtained from a solid mixture of **V** and **W**. (2 Marks)

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17. Use the bond energies given below to calculate the heat of reaction for; (3 marks)



Bond	Energy (Kj/Mol)
H – H	435
Cl – Cl	243
H – Cl	431

18. The PH of a soil sample was found to be 5.7. An agricultural officer recommended addition of lime.

(a) State **two** functions of the lime. (2 Marks)

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(b) Give the name of the process applied in (a) above. (1 mark)

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19. The electronic configuration of ions  $X^{2+}$  is 2.8 while that of ion  $Y^-$  is 2.8.8.

(a) Write down the electron arrangement of the atoms of X and Y (2 Marks)

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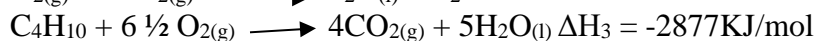
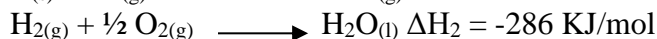
(b) Compare the atomic radii of the two elements. (1 Mark)

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(c) Give the name of the chemical family to which element X belongs (1 Mark)

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20. Use the information below to answer the questions that follow.



(a) Calculate the molar enthalpy of formation of butane ( $C_4H_{10}$ ) from its elements in their normal states.

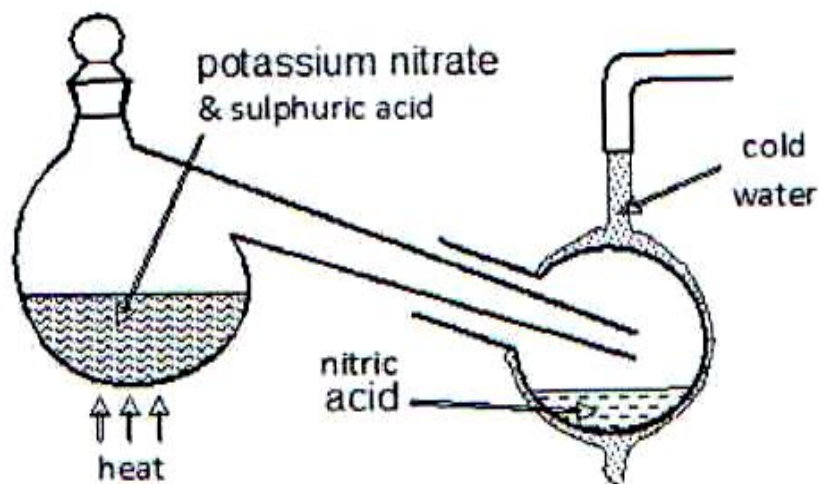
(3mks)

21. (a) (i) A student found a colourless liquid in the laboratory which he suspected to be water. Describe a chemical test he could have performed to confirm that the liquid is water. (2 Marks)

(ii) What other test could he have done to prove that the liquid is pure water?(1 Mark)

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22. The diagram below shows that the set-up that was used to prepare and collect a sample of nitric acid



(a) Give a reason why it is possible to separate nitric acid from sulphuric acid in the set-up.  
(1 Mark)

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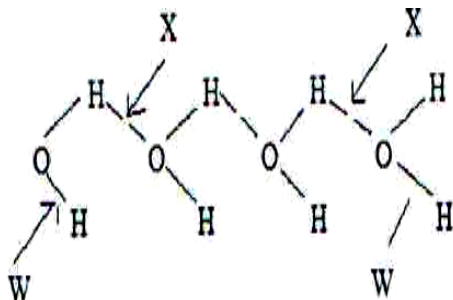
(b) Name another substance that can be used instead of potassium nitrate.(1 Mark)

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(c) Give one use of nitric acid.(1mark)

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23. The structure of water molecules can be represented as shown below.



(i) Name the bond type represented by letter X and W.

*(1 Mark)*

(ii) Relative molecular mass of methane and water are almost similar, however the boiling of water is  $100^{\circ}\text{C}$  while that of methane is  $-161^{\circ}\text{C}$ . Explain.

*(1 Mark)*

24. Diamond and graphite are allotropes of carbon. In terms of structure and bonding, explain why?

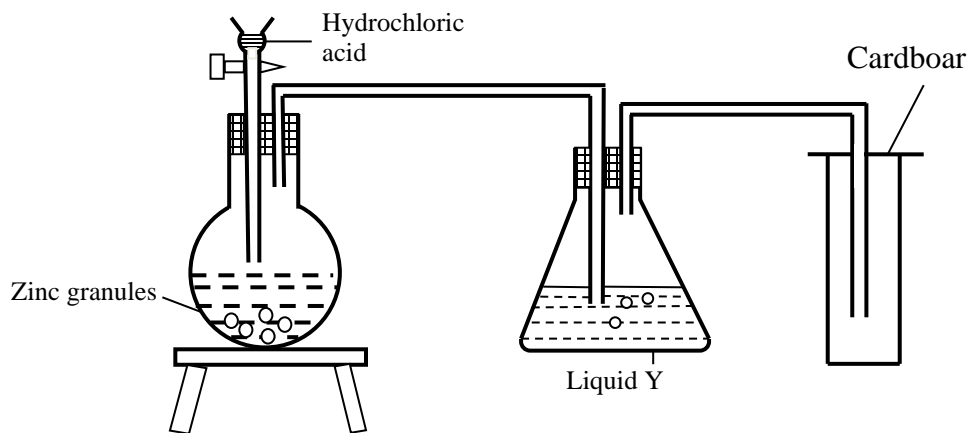
(i) Diamond is used in drilling of hard rocks.

*(1 Mark)*

(ii) Graphite is a lubricant.

*(1 Mark)*

25. The set up was used to prepare dry hydrogen gas. Study it and answer the questions that follow.



(i) Is set-up used to prepare the gas correct? Give reason.

(1 Mark)

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(ii) What would be liquid Y? (1 mark)

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(iii) Give two physical properties of hydrogen gas

(1 Mark)

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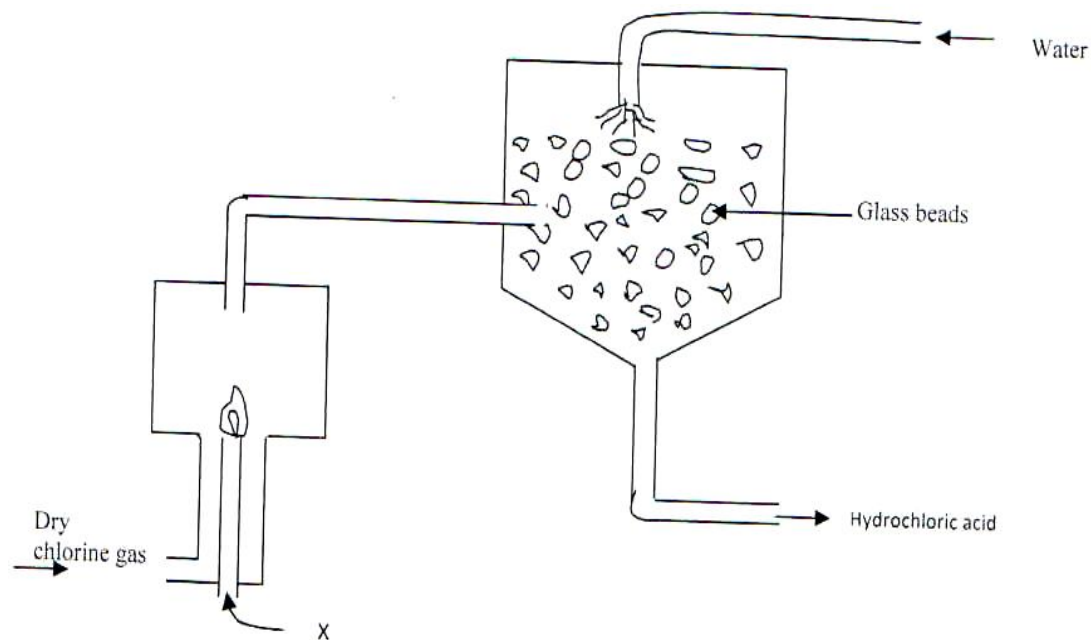
26. Given element W has atomic number 14 and consists of isotopes as shown below.

Isotope	A	B	C
Isotope mass	28	29	30

Percentage abundance                      92.2   4.7   3.1  
Determine the relative atomic mass of W

(2 Marks)

27. The diagram below represents a set up used for the large scale manufacture of hydrochloric acid.



(a) Name substance X

(1Mark)

(b) What is the purpose of the glass beads?

(1 Mark)

(c) Give one use of hydrochloric acid

(1Mark)

28. A mixture contains Iron (III) Chloride, calcium chloride and iron filings. Describe how one can separate and recover the substances in the mixture.(3marks)

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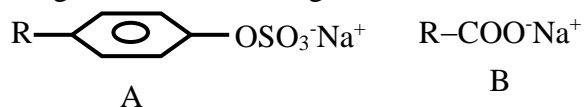
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29. The structure below represents two cleansing agents A and B. Which cleansing agent would be suitable for washing in water containing calcium chloride? Give a reason.(2marks)



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